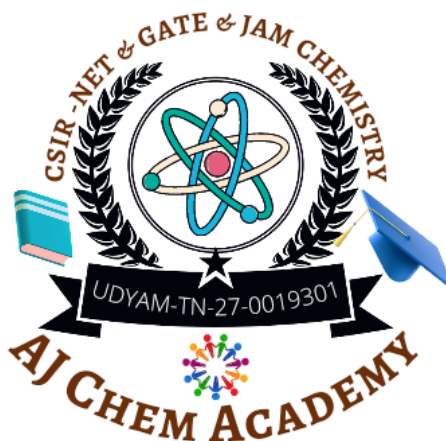


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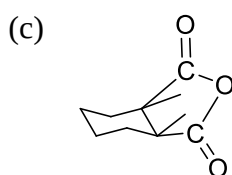
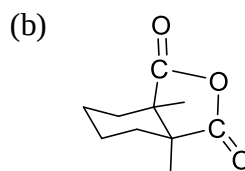
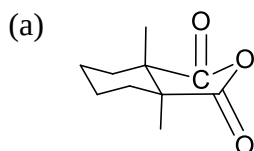
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Q.1 – Q.75 MCQ, carry Four marks each. (for each wrong answer: -1).

You are required to Answer All Questions.

1. Reaction of **diborane** with **excess ammonia** at **low temperature** gives initially
 (a) $B_2H_6 \cdot NH_3$ (b) Borazole (c) $[BH_2(NH_3)_2]^+ [BH_4]^-$ (d) $B_2N_4H_{10}$
2. Which of the following isomer of **1,2-cyclohexanedicarboxylic acid** form **anhydride** on heating?



(d) Reaction not possible

3. Arrange the **increasing order of charges (molar conductivity)** of the compounds used in **Werner coordination theory**



P



Q



R



S

Choose the correct answer from the options given below:

- (a) P, Q, R, S (b) S, R, Q, P (c) Q, P, S, R (d) R, Q, S, P
4. Match List-I with List-II

	List-I (Energies)		List-II (Values)
P.	Energy of ground state of He^+	I.	+6.04 eV
Q.	Potential energy of I orbit of H-atom	II.	-27.2 eV
R.	Kinetic Energy of II excited state of He^+	III.	+54.4 eV
S.	Ionization potential of He^+	IV.	-54.4 eV

Choose the correct answer from the options given below:

	P		Q		R		S																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																
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5. Calculate the **wavelength (in nanometer)** associated with a proton moving at $1.0 \times 10^3 \text{ ms}^{-1}$,
 (a) 0.032 nm (b) 0.40 nm (c) 2.5 nm (d) 14.0 nm
6. **Number of distinct 1H -NMR signals** observed in case of **acetone** and **ethyl methyl**

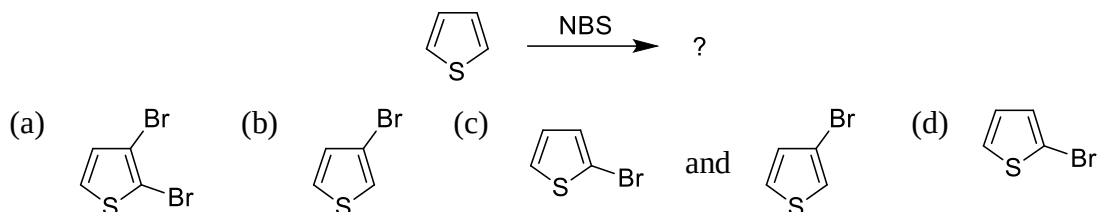
ketone are

- (a) 1 and 3 (b) 2 and 3 (c) 2 and 5 (d) 1 and 2

7. Oxidation state of each Cl in CaOCl_2 is/are

- (a) 0 (b) +1 (c) -1 (d) +1, -1

8. Identify the product in the following chemical reaction?



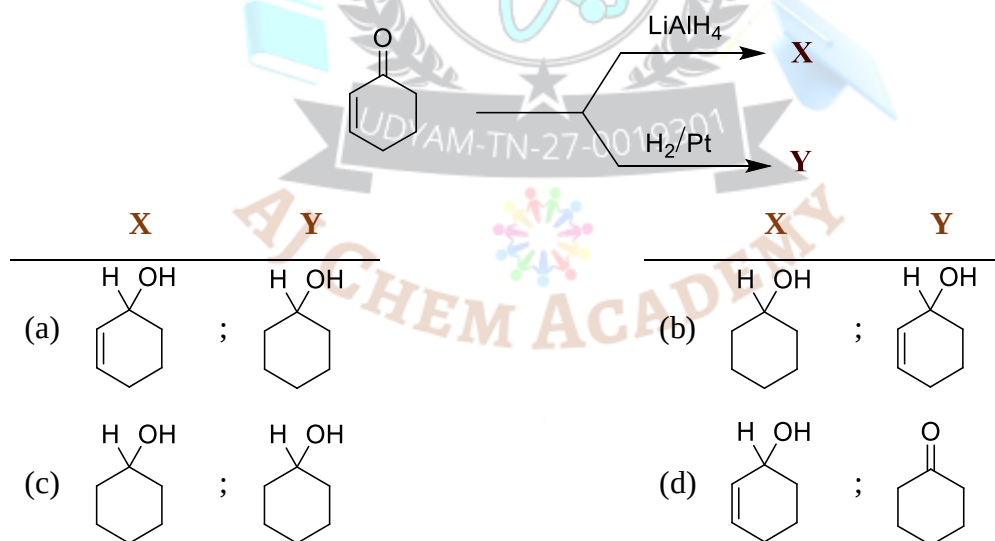
9. The unit of van der Waal's constant 'a' is

- (a) $\text{dm}^3 \text{mol}^{-1}$ (b) $\text{dm}^6 \text{atm mol}^{-2}$ (c) $\text{dm}^6 \text{mol}^{-1}$ (d) $\text{dm}^2 \text{atm mol}^{-1}$

10. The number of degrees of freedom for liquid water and water vapour in equilibrium at a pressure of 1 atm is

- (a) 0 (b) 1 (c) 2 (d) 3

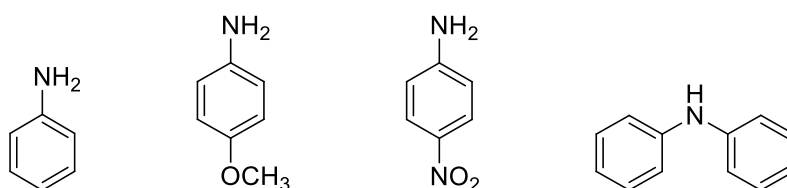
11. In the following reaction, identify the product X and Y



12. An electron is confined to move in a one-dimensional box of length 1 Å. Its energy in the first excited state is approximately

- (a) 150.4 eV (b) 112.8 eV (c) 37.6 eV (d) 342.0 eV

13. To arrange the following compounds in the order of basicity:



I

II

III

IV

(a) II > I > III > IV

(b) IV > III > II > I

(c) I > II > III > IV

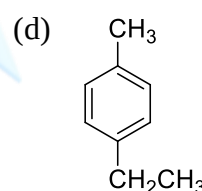
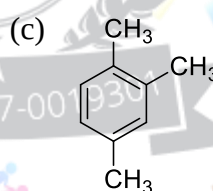
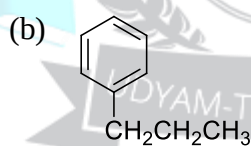
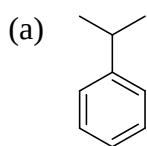
(d) III > IV > I > II

14. Match List-I with List-II

	List I (Concentration)		List-II (Units)
P.	Molarity	I.	Mol
Q.	Mole fraction	II.	unitless
R.	Mole	III.	mol L ⁻¹
S.	Molality	IV.	mol kg ⁻¹

Choose the correct answer from the options given below:

P	Q	R	S	P	Q	R	S
(a) I ; II ; IV ; III	(b) III ; II ; I ; IV						
(c) II ; I ; III ; IV	(d) II ; III ; I ; IV						

15. The structure of the compound having molecular formula C_9H_{12} showing NMR peaks δ 7.1, 2.2, 1.5 and 0.9 ppm is

16. For a zero order reaction

P. The rate of the reaction is independent of reactant concentration

Q. The rate constant is independent of temperature

R. The rate constant of the reaction is independent of reactant concentration

S. Thermal decomposition of HI on gold surface is a zero order reaction

Choose the correct feature(s) from the options given below:

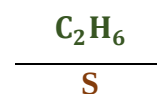
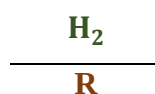
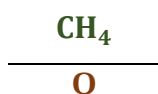
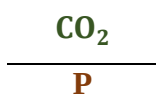
(a) P only

(b) P and Q only

(c) P, R and S only

(d) Q and R only

17. The increasing order of number of vibrational degrees of freedom from the following



Follows the order:

(a) P, Q, R, S

(b) P, S, R, Q

(c) Q, P, R, S

(d) R, P, Q, S

18. Which of the following **vibrational mode** shows **no IR absorption bands**

- (a) symmetric CO_2 stretch (b) antisymmetric CO_2 stretch
(c) symmetric $\text{O}=\text{C}=\text{S}$ stretch (d) $\text{H}_3\text{C}-\text{C}\equiv\text{CH}$ stretch

19. Match List-I with List-II

	List-I (Quantity)		List-II (Conversion)
P.	88 g of CO_2	I.	2 mol
Q.	6.023×10^{23} molecules of H_2O	II.	1 mol
R.	96 g of O_2	III.	6.023×10^{23} molecules
S.	1 mol of any gas	IV.	3 mol

Choose the correct answer from the options given below:

P	Q	R	S	P	Q	R	S
(a) I ; II ; IV ; III	(b) III ; II ; I ; IV						
(c) II ; I ; III ; IV	(d) II ; III ; I ; IV						

20. Which one of the following is the **weakest base**?

- (a) $\text{Ca}(\text{OH})_2$ (b) KOH (c) LiOH (d) $\text{Sr}(\text{OH})_2$

21. **Electronegativity** of the following elements **increases** in the order

Carbon	Nitrogen	Silicon	Phosphorus
P	Q	R	S

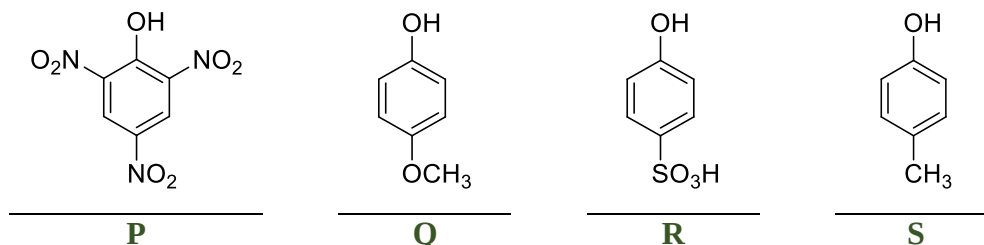
Choose the correct answer from the options given below:

- (a) R, S, P, Q (b) R, P, S, Q (c) Q, P, S, R (d) R, Q, S, P

22. **Mg-ATP complex** is a substrate for

- (a) Kinase (b) Catalase (c) Helicase (d) Hydrolyses

23. Arrange the following compounds **P, Q, R** and **S** in **decreasing order of acidic strength**



Choose the correct answer from the options given below:

- (a) $Q > S > R > P$ (b) $P > R > S > Q$
(c) $Q > P > S > R$ (d) $Q > R > P > S$

24. In $\text{S}_{\text{N}}1$ and $\text{S}_{\text{N}}2$ reactions

P. S_N1 is a unimolecular reaction with first order kinetics while S_N2 reaction is bimolecular reaction with second order kinetics

Q. In S_N2 , the reaction proceeds through the formation of carbocation while S_N1 does not

R. S_N2 is a stereospecific reaction. The product formed with inversion of configuration only

S. S_N1 reaction is favoured in the presence of weak base or poor nucleophile

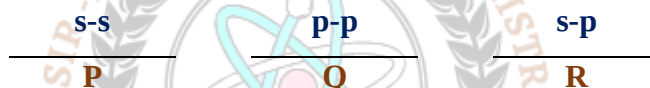
Choose the correct answer from the options given below:

(a) P, Q and S only (b) P, Q and R only (c) P, R and S only (d) Q, R and S only

25. The expression relating **molality (m)** and **mole fraction (χ_2)** of solute in a solution is

(a) $\chi_2 = \frac{mM_1}{1+mM_1}$ (b) $\chi_2 = \frac{mM_1}{1-mM_1}$ (c) $\chi_2 = \frac{1+mM_1}{mM_1}$ (d) $\chi_2 = \frac{1-mM_1}{mM_1}$

26. The **decreasing strength** of bond formed by **overlap** of the following



Follows the order:

(a) R, P, Q (b) Q, P, R (c) P, Q, R (d) P, R, Q

27. Consider the following statements:

P. 1, 2-addition reaction occurs faster as compared to 1, 4-addition reaction but 1, 4-addition product is more stable

Q. Formation of 1, 2-addition product is kinetic or rate controlled

R. Formation of 1, 4-addition product is thermodynamic or equilibrium controlled

S. At low temperature the formation of 1, 2-addition product from allyl cation is a reversible reaction

Choose the correct answer from the options given below:

(a) P, Q and S only (b) P, Q and R only (c) P, Q, R and S (d) Q, R and S only

28. Match List-I with List-II

	List-I (Thermodynamic properties)		List-II (Consequences in Reaction)
P.	$\Delta H = -ve, \Delta S = -ve, \Delta G = -ve$	I.	Non-Spontaneous at High Temperature
Q.	$\Delta H = -ve, \Delta S = -ve, \Delta G = +ve$	II.	Non-Spontaneous at Low Temperature
R.	$\Delta H = +ve, \Delta S = +ve, \Delta G = +ve$	III.	Spontaneous at Low Temperature

S. $\Delta H = +ve, \Delta S = +ve, \Delta G = -ve$ **IV.** Spontaneous at High Temperature

Choose the correct answer from the options given below:

P	Q	R	S	P	Q	R	S
(a) II ; III ; I ; IV	(b) III ; I ; II ; IV						
(c) III ; II ; IV ; I	(d) II ; IV ; I ; III						

29. Which of the following species has the **highest electron affinity**?

- (a) F^- (b) Cl^- (c) O^- (d) Na^-

30. Which of the following oxides is **not expected to react with sodium hydroxide**?

- (a) CaO (b) SiO_2 (c) BeO (d) B_2O_3

31. The correct order of the stability of different **conformation of cyclohexane** is

chair	boat	twist boat	Half chair
P	Q	R	S
(a) $P > Q > R > S$	(b) $P > R > Q > S$		
(c) $P > S > R > Q$	(d) $S > R > Q > P$		

32. Among the given **characteristics of viscosity of liquids**

P. Greater the viscosity, more slowly the liquid flows

Q. Glass is an extremely viscous liquid

R. Viscosity of liquid increases as the temperature rises

S. Fluidity is the reciprocal of viscosity

Choose the correct characteristics from the options given below:

- (a) P, Q and S only (b) P and R only (c) Q and R only (d) P, Q, R and S

33. Identify the structure of the **alkaloid** having the following characteristics;

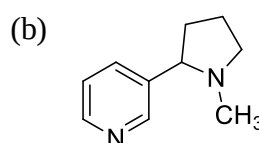
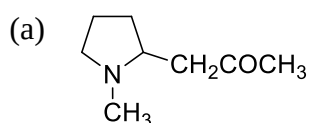
P. It is coca alkaloid having molecular formula $C_8H_{15}NO$

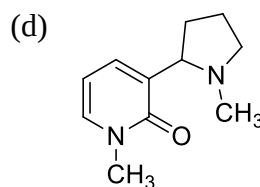
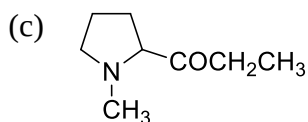
Q. As the free base, it rapidly racemizes

R. Its reactions show the presence of keto group and a tertiary nitrogen atom

S. Synthesized by condensing γ -methylaminobutyraldehyde with ethylacetoacetate in buffered solution at $pH = 7$

Choose the correct answer from the options given below:





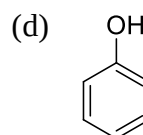
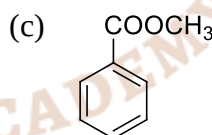
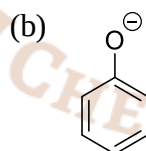
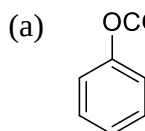
34. Which of the following statement is **not correct**

- (a) Viscosity of ethanol is smaller than that of glycol
- (b) Viscosity of liquids increases with increase of pressure
- (c) The variation of viscosity of liquids with temperature is given by $\eta = Ae^{-E/RT}$
- (d) The volume of liquid flowing in time "t" through a pipe is given by Poisseuille equation

35. Which of the following regarding the **shapes of d-orbitals** is **not correct**

- (a) d_z^2 is a dumb-bell shaped curve symmetric about z-axis with a ring like collar in xy plane
- (b) $d_{x^2-y^2}$ is double dumb-bell shaped with signs of the lobes on y-axis will always be Positive
- (c) d_{xy} is double dumb-bell shaped
- (d) A 3d orbital would have no radial node

36. Which of the following compounds will have the **highest order of their reactivity** towards **electrophilic substitution reaction**:



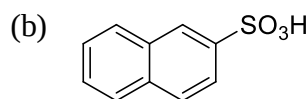
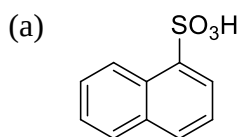
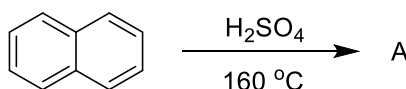
37. **Stability of the species Li_2 , Li_2^- and Li_2^+** increases in the order of:

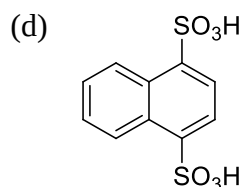
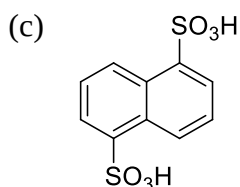
- (a) $Li_2^- < Li_2 < Li_2^+$
- (b) $Li_2 < Li_2^+ < Li_2^-$
- (c) $Li_2^- < Li_2^+ < Li_2$
- (d) $Li_2 < Li_2^- < Li_2^+$

38. The **basic structural unit of silicates** is

- (a) $[SiO_3]^{2-}$
- (b) $[SiO_4]^{2-}$
- (c) $[SiO]^-$
- (d) $[SiO_4]^{4-}$

39. The **product-A** in the following reaction is:





40. Match List-I with List-II

	Ion		Shape		P	Q	R	S			
P	$[\text{ICl}_2]^-$	I	V-shape	(a)	II	;	III	;	I	;	IV
Q	$[\text{NH}_2]^-$	II	Linear	(b)	III	;	II	;	I	;	IV
R	$[\text{NH}_4]^+$	III	Tetrahedral	(c)	II	;	I	;	III	;	IV
S	$[\text{PtCl}_4]^{2-}$	IV	Square planar	(d)	III	;	IV	;	I	;	II

41. Which one is **not** an electrophile?

- (a) NO^+ (b) BF_3 (c) CO_2 (d) $[\text{NH}_4]^+$

42. What is the **oxidation state** of Pd in $[\text{Pd}(\text{OAc})_2]$?

- (a) 0 (b) 1 (c) 2 (d) 3

43. What is the **EAN** of $[\text{Co}(\text{NH}_3)_6]^{3+}$?

- (a) 36 (b) 38 (c) 35 (d) 54

44. **Cytochrome-C** is a/an

- (a) Electron donor (b) Electron acceptor (c) Neutral (d) Negatively charged

45. The **atomic masses** of $^1\text{H} = 1.673 \times 10^{-27} \text{ kg}$ and $^{35}\text{Cl} = 58.06 \times 10^{-27} \text{ kg}$. The **reduced mass** of HCl is

- (a) $162.6 \times 10^{-27} \text{ kg}$ (b) $16.26 \times 10^{-27} \text{ kg}$
(c) $1.626 \times 10^{-27} \text{ kg}$ (d) $1626 \times 10^{-27} \text{ kg}$

46. The **ionic strength** of a solution which is **0.1 molal** in **NaCl** and **0.01 molal** in **calcium chloride** is (assuming complete dissociation)

- (a) 0.13 molal (b) 0.26 molal (c) 0.11 molal (d) 0.056 molal

47. Consider the following statements with reference to **Citral**:

P. Citral is a liquid which has a smell of lemon

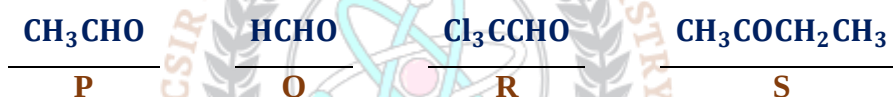
Q. Oxidation of citral with alkaline permanganate followed by chromic acid gives acetone, oxalic acid and laevulic acid

R. It also forms oxime

S. It shows two geometrical isomers: trans isomer known as neral and cis isomer known as geranial

Choose the correct answer from the options given below:

- (a) P, Q, R and S (b) P, Q and R only (c) P, Q and S only (d) Q, R and S only
48. The expression of **average speed** (v_{av}) of molecules of a gas is given by
- (a) $\left(\frac{8RT}{\pi m}\right)^{1/2}$ (b) $\left(\frac{8RT}{\pi M}\right)^{1/2}$ (c) $\left(\frac{3RT}{\pi m}\right)^{1/2}$ (d) $\left(\frac{3RT}{m}\right)^{1/2}$
49. The **magnitude of angular momentum** of an electron occupying **3s atomic orbital** is
- (a) $\sqrt{2} \hbar$ (b) $\sqrt{6} \hbar$ (c) $\sqrt{3} \hbar$ (d) 0
50. In the **conductometric titration of hydrochloric acid against ammonium hydroxide**
- (a) The conductance increases till the end point with the volume of NH_4OH added
 (b) The conductance remains more or less constant till the end point
 (c) The conductance remains more or less constant after the end point
 (d) The conductance increases after the end point with the volume of NH_4OH added
51. Arrange the following **carbonyl compounds** in the **decreasing order** of their reactivity towards **nucleophilic addition reaction**



Choose the correct answer from the options given below:

- (a) $P > Q > R > S$ (b) $Q > P > S > R$
 (c) $Q > R > P > S$ (d) $R > Q > P > S$
52. The **probability of finding the electron in the p_x orbital** is
- P. Zero at the nucleus**
Q. Maximum on two opposite sides of the nucleus along x-axis
R. Zero on z-axis
S. Same on all sides around the nucleus
- Choose the correct answer from the options given below:
- (a) P, Q and S only (b) P, Q and R only (c) P, Q, R and S (d) Q, R and S only
53. Match List-I with List-II

	List-I		List-II
P.		I.	Oxepane
Q.		II.	Oxolane
R.		III.	Oxetane

S.



IV.

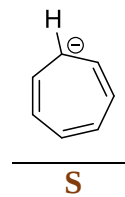
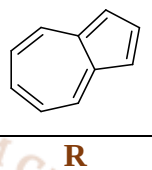
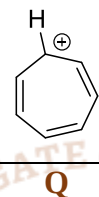
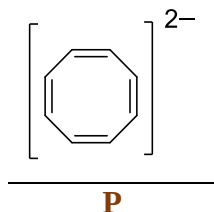
Oxirane

Choose the correct answer from the options given below:

P	Q	R	S
(a) III ; IV ; II ; I			
(c) I ; II ; IV ; III			

P	Q	R	S
(b) III ; II ; I ; IV			
(d) I ; III ; IV ; II			

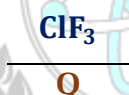
54. Which of the following molecule is **aromatic**?



Choose the correct answer from the options given below:

- (a) Q and R only (b) P, Q and R only (c) P, R and S only (d) Q, R and S only

55. The compound(s) with **two lone pairs of electrons** on the central atom is (are)



Choose the correct answer from the options given below:

- (a) P, Q and S only (b) Q and R only (c) Q, R and S only (d) P, R and S only

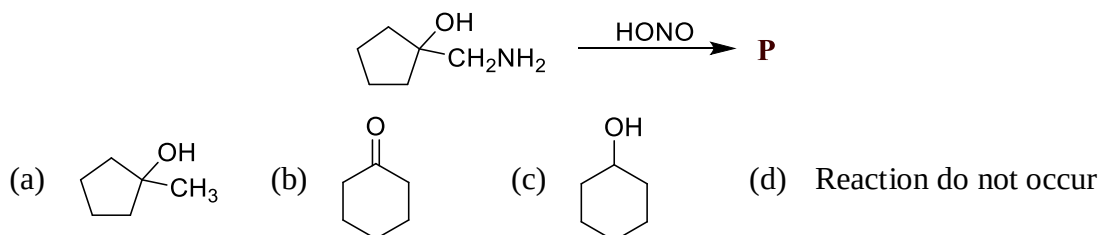
56. The **molar ionic conductance at infinite dilution** for NaOH, NaCl and BaCl₂ are **248.1×10^{-4} , 126.5×10^{-4} and $280.0 \times 10^{-4} \text{ S m}^2 \text{ mol}^{-1}$** respectively. The Λ_m^0 for **Ba(OH)₂** is

- (a) $523.2 \times 10^{-4} \text{ S m}^2 \text{ mol}^{-1}$ (b) $5.232 \times 10^{-4} \text{ S m}^2 \text{ mol}^{-1}$
 (c) $52.3 \times 10^{-4} \text{ S m}^2 \text{ mol}^{-1}$ (d) $50.5 \times 10^{-4} \text{ S m}^2 \text{ mol}^{-1}$

57. The **Miller indices of crystal planes** which cut through the crystal axes at **(2a, 3b, c)** are

- (a) (326) (b) (231) (c) (132) (d) ($3\bar{2}\bar{2}$)

58. Identify the **product-P** in the following chemical reaction



59. The **increasing order of molarity** of 25 gm each of

NaOH

P

LiOH

Q

KOH

R

Al(OH)₃

S

B(OH)₃

T

In same volume of water is

- (a) S < T < R < P < Q (b) S < R < Q < P < T
 (c) Q < P < R < T < S (d) Q < R < S < P < T

60. Match List-I with List-II

	List-I		List-II
P.	$\text{Ph}_2\text{C}=\text{NOH} \rightarrow \text{PhCONHPh}$	I.	Knoevenagel condensation
Q.	$\text{CH}_3\text{COCH}_3 \rightarrow \text{CH}_3\text{COOCH}_3$	II.	Claisen-Schmidt reaction
R.	$\text{PhCHO} \rightarrow \text{PhCH}=\text{CHCOOH}$	III.	Beckmann rearrangement
S.	$\text{PhCHO} \rightarrow \text{PhCH}(\text{OH})\text{CH}_2\text{COCH}_3$	IV.	Baeyer Villiger oxidation

Choose the correct answer from the options given below:

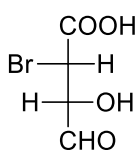
P	Q	R	S	P	Q	R	S
(a) III ; IV ; I ; II				(b) II ; III ; I ; IV			
(c) I ; II ; IV ; III				(d) III ; I ; II ; IV			

61. From the following half-cell reactions

The standard potential E° of the cell, $\text{Zn}, \text{Zn}^{2+}(1\text{M}) \mid \text{I}^{-}(1\text{M}), \text{CuI}; \text{Cu}$, will be

- (a) -0.42 V (b) $+1.10 \text{ V}$ (c) $+0.42 \text{ V}$ (d) $+0.59 \text{ V}$

62. Assign the R and S configuration to the following molecule



- (a) 2S, 3R (b) 2R, 3S
 (c) 2S, 3S (d) 2R, 3R

63. Which of these will **not** be oxidised by ozone?

- (a) KI (b) FeSO_4 (c) KMnO_4 (d) K_2MnO_4

64. Europium (Eu) resembles Calcium (Ca) in the following ways:

P. Both are diamagnetic

Q. Insolubility of their sulphates and carbonates in water

R. Solubility of these metals in liquid NH_3

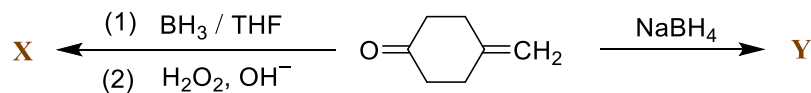
S. Insolubility of their dichlorides in strong HCl

Choose the correct answer from the options given below:

- (a) P, Q and S only (b) P, Q and R only (c) P, Q, R and S (d) Q, R and S only

65. The plot of $\log k$ versus $1/T$ of a reaction is linear with a
- (a) Positive slope and zero intercept (b) Positive slope and nonzero intercept
(c) Negative slope and zero intercept (d) Negative slope and nonzero intercept

66. Give the structure of **X** and **Y**



- (a) O=C1C=CC=CC=C1 ; OCC1C=CC=CC=C1
 (b) OCC1C=CC=CC=C1 ; O=C1C=CC=CC=C1
 (c) O=C1C=CC=CC=C1 ; OCC1C=CC=CC=C1
 (d) O=C1C=CC=CC=C1 ; O=C1C=CC=CC=C1

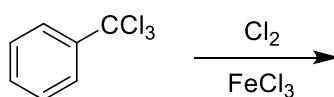
67. SF_2 , SF_4 and SF_6 have the hybridisation at sulphur atom respectively as
- (a) $\text{sp}^2, \text{sp}^3, \text{sp}^3\text{d}^2$ (b) $\text{sp}, \text{sp}^3, \text{sp}^3\text{d}^2$ (c) $\text{sp}^3, \text{sp}^3\text{d}, \text{sp}^3\text{d}^2$ (d) $\text{sp}^3, \text{sp}^2, \text{d}^2\text{sp}^3$
68. The Lanthanide contraction from Ce to Lu is
- (a) $0.20 \text{ \AA}$ (b) $0.10 \text{ \AA}$ (c) $0.30 \text{ \AA}$ (d) $0.25 \text{ \AA}$
69. Match List-I with List-II

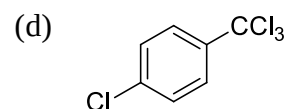
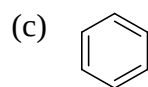
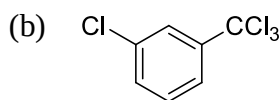
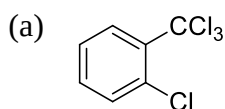
	List-I		List-II
	Metal/Compounds		Catalytic properties
P	TiCl_3	I	Adams catalyst
Q	FeSO_4	II	Repp synthesis
R	Pt/PtO	III	Used as the Ziegler-Natta catalyst
S	Ni	IV	Used as Fenton's reagent

Choose the correct answer from the options given below:

- | | | | | | | | |
|-----------------------|-----------------------|----------|----------|----------|----------|----------|----------|
| P | Q | R | S | P | Q | R | S |
| (a) I ; II ; III ; IV | (b) I ; III ; II ; IV | | | | | | |
| (c) I ; II ; IV ; III | (d) III ; IV ; I ; II | | | | | | |

70. Find the major product in the following reaction





71. Match List-I with List-II

	Ions		Magnetic moment (BM)
P.	Co ²⁺	I.	1.73
Q.	Mn ²⁺	II.	3.87
R.	Cr ²⁺	III.	4.90
S.	Cu ²⁺	IV.	5.92

Choose the correct answer from the options given below:

	P	Q	R	S		P	Q	R	S
(a)	II	;	IV	;	III	;	I		
(b)	I	;	III	;	II	;	IV		
(c)	I	;	II	;	IV	;	III		
(d)	III	;	IV	;	I	;	II		

72. Match List-I with List-II

	Formula		Structure
P.	XeO ₄	I.	Pyramidal
Q.	XeO ₂ F ₂	II.	Tetrahedral
R.	XeF ₄	III.	Trigonal bipyramidal
S.	XeO ₃	IV.	Square planar

Choose the correct answer from the options given below:

P	Q	R	S	P	Q	R	S
(a) II	;	III	;	IV	;	I	
(c) I	;	II	;	IV	;	III	

P	Q	R	S	P	Q	R	S
(b) II	;	IV	;	III	;	I	
(d) I	;	III	;	IV	;	II	

73. Iodoform reaction is given by

P. All methyl ketones with CH₃CO group

Q. Acetaldehyde

R. All secondary alcohols with CH₃CH(OH) group

S. All primary alcohols

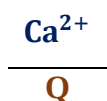
Choose the correct answer from the options given below:

(a) P, Q and S only (b) P, Q and R only (c) P, Q, R and S (d) Q, R and S only

74. The reduction potential of hydrogen half cell will be negative if

(a) p(H₂) = 1 bar and [H⁺] = 1M(b) p(H₂) = 1 bar and [H⁺] = 2M(c) p(H₂) = 2 bar and [H⁺] = 1M(d) p(H₂) = 2 bar and [H⁺] = 2M

75. The correct order of **decreasing ionic radii** among the following **isoelectronic species** is



(a) $\text{P} > \text{Q} > \text{R} > \text{S}$

(b) $\text{P} > \text{R} > \text{Q} > \text{S}$

(c) $\text{S} > \text{Q} > \text{R} > \text{P}$

(d) $\text{S} > \text{R} > \text{P} > \text{Q}$

Answer Key

Q.No	Ans
1.	c
2.	b
3.	b
4.	a
5.	b
6.	a
7.	d
8.	d
9.	b
10.	a
11.	d
12.	a
13.	a
14.	b
15.	b

Q.No	Ans
16.	c
17.	d
18.	a
19.	a
20.	c
21.	a
22.	a
23.	b
24.	c
25.	a
26.	d
27.	b
28.	b
29.	c
30.	a

Q.No	Ans
31.	b
32.	a
33.	a
34.	c
35.	b
36.	b
37.	c
38.	d
39.	b
40.	c
41.	d
42.	c
43.	a
44.	b
45.	c

Q.No	Ans
46.	a
47.	b
48.	b
49.	d
50.	c
51.	c
52.	b
53.	d
54.	b
55.	b
56.	a
57.	a
58.	b
59.	a
60.	a

Q.No	Ans
61.	d
62.	a
63.	c
64.	d
65.	d
66.	a
67.	c
68.	b
69.	d
70.	b
71.	a
72.	a
73.	b
74.	c
75.	d

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